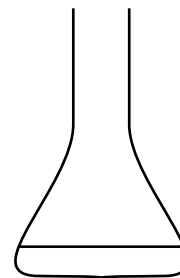


## Aim

To boil water at less than 100 °C by lowering the air pressure

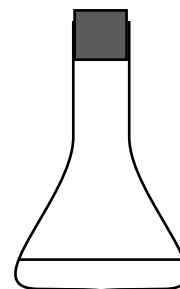
## Method

1. We boiled 1 cm of water in a conical flask.



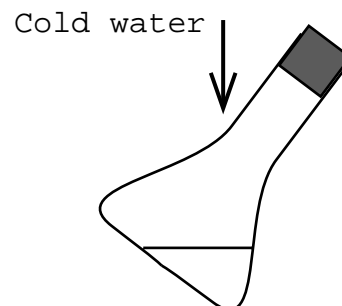
Heat

2. We removed the Bunsen and then quickly fitted a bung.



Do not heat

3. We then ran the flask under cold water.



## Results

The water...

The bung was ...

## Conclusion

This shows the boiling point of a liquid depends on the ..... (1 word).

Running cold water over the conical flask *lowered/raised* the air pressure (because the air inside the flask contracts). The water boils because there is *less/more* air pressure to keep it as a liquid.

This is the same as water boiling at *lower/higher* temperatures as you climb a mountain.

The bung is hard to remove because

## Answers

R. boiled again when the cold water was running over the flask /. hard to remove.

C. pressure: / lowered, less, higher / more air pressure on the outside than the inside or pulling against the weight of the atmosphere.